

INTERNATIONAL INDIAN SCHOOL RIYADH
CLASS IX WORKSHEET 2017-18 CHEMISTRY

1. Name the subatomic particles which are same in two isotopic species of an atom
2. What would be the valence electrons in an atom X having atomic number 16 ?
3. An element X has a mass number 4 and atomic number 2. Find the valency of X.
3. Write the main drawback of Rutherford's model of atom.
4. Give reason:
 - (a) Noble gases show least reactivity
 - (b) Nucleus of an atom is heavy and positively charged.
5. Define the terms isotopes and isobars giving one examples for each.
6. List the postulates of Bohr's model of atom.
7. Give the schematic atomic structure of Chlorine atom and Chlorine ion.
8. Explain Bohr and Bury rules for distribution of electrons into different shells
9. Write the electronic configuration of an element X with atomic number 15.
10. An isotope of Chlorine is represented as ${}_{17}^{37}\text{Cl}$. Calculate the number of electrons , protons and neutrons.
11. Derive the valency of an atom having atomic number 5. Name the element which has no neutron in its atom.
- 12 .What will be the composition of the nucleus of the atom of an element with atomic number 19 and mass number 39. Write its electronic configuration.
13. Which has more electrons Na or Na^+ . Why?

14. Name the scientist who discovered neutrons. What is the charge and mass of a neutron? Where is the neutron located in atom?

15. Why is the atomic mass of Chlorine taken as 35.5 u and not a whole number . Explain.

16. What are canal rays ? Who discovered them? What is the charge and mass of a canal ray